

chapter 10 chemical quantities pdf

What next? Now substitute the unit grams for atomic mass units to find the molar mass of SO₃. 1 mol of SO₃ has a mass of 80.1 g. This is the mass of 6.02 x 10²³ molecules of SO₃

Chapter 10: Chemical Quantities - Dr Collings' Science Classes

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Chemical Quantities THE MOLE AND QUANTIFYING MATTER 10.1 The Mole: A Measurement of Matter Essential Understanding The mole represents a large number of very small particles. Reading Strategy Frayer Model The Frayer Model is a vocabulary development tool. The center of the diagram shows the concept being defined, while the quadrants around the concept are used for providing the details. Use ...

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of all elements a. are the same. Guided Notes Chapter 10: Chemical Quantities Name 10.1 The ... guided notes chapter 10: chemical quantities name_____ 10.1 the mole: a measurement of matter 1.

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CHAPTER 10: Chemical Quantities BASICS: The basic unit that is used to determine the amount of a chemical substance is called a mole A mole(mol) of a substance is equivalent to 6.02 x 10²³ particles of that substance The mole was founded by a scientist named Avagadro, and he decided to use the number 6.02 x 10²³ because thats how many atoms are in 12.01 grams of carbon A very ...

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